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are not available in a pure state. For example, a mixture of reactants may contain a wide variety of different species. Most of the reaction products are not always suitable for recovery or separation. In these cases, catalytic enhancement of the reaction may provide an opportunity to recover pure products. After the reaction, catalysts may be used to speed the reaction, enhance the reactivity of the reaction, modify the products formed, or alter a rate constant or selectivity of a particular product. One way of understanding the functioning of a catalyst may be to consider it as a species that lowers the activation energy or the transition state of a reaction. A catalyst raises the substrate-to-product rate transition state and reduces the Gibbs free energy barrier separating the reactants and products of the reaction. In more general terms, an active catalyst lowers the free energy barrier associated with the reaction. The nature of the active catalyst may be related to the chemical and physical

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properties of the catalyst. Certain types of catalysts may form additional products, called by-products, in addition to the desired products. These by-products may also lower the energy barrier. To avoid a buildup of by-products, the desired reaction has to be terminated to allow the catalyst to recover its original high activity. Thus, catalysis requires a system that allows for the regeneration of the catalyst and the separation of the various products formed. As an example, catalysts may be used to improve the yield of organic compounds such as medicines from natural products. In this case, a catalyst may convert certain unstable molecules into more stable compounds. This conversion process may involve multiple steps, such as depolymerization, hydrolysis, and dec

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